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## Using elementary reactions to model growth processes of polyaromatic hydrocarbons under pyrolysis conditions of light feedstocks

K. Hemelsoet<sup>a</sup>, V. Van Speybroeck<sup>a\*</sup>, K.M. Van Geem<sup>b</sup>, G.B. Marin<sup>b</sup> and M. Waroquier<sup>a</sup>

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Density functional theory results are presented for elementary steps leading to coke growth within a steam cracking unit. The discussed pathway starts from toluene and ultimately, 1-methylnaphthalene is formed. In order to find the rate determining step for coke formation, the pseudo first-order rate coefficients of the various steps are compared taking into account the concentrations of diverse coke precursors. The influence of the polyaromatic environment is studied for a large set of methylated polycyclic aromatic molecules, by means of carbon–hydrogen bond dissociation enthalpy values. Subsequent hydrogen abstraction reactions at the methylated polyaromatics, by a methyl radical, are also examined. The abstraction is found to preferentially occur at the larger systems and is in general faster compared to abstractions at the analogous non-methylated species.

**Keywords:** DFT calculations; polyaromatic hydrocarbons; bond dissociation enthalpy; carbon-centered radicals; hydrogen abstraction

### 1. Introduction

Various free radical mechanisms leading to the growth of polyaromatic species have in recent years intensively been examined using *ab initio* modelling techniques [1–5]. In this light, the presenting authors have provided important contributions focussing on coke growth within a steam cracking unit [6–11]. Steam cracking of large hydrocarbons remains at present the most important production source of light olefins such as ethylene and propylene. Simultaneously with the cracking reactions, the growth of a coke layer at the inner wall of the reactor tube is taking place. The presence of this layer reduces the efficiency of the cracking unit substantially and its growth is thus highly undesirable. In practice, thousands of reactions are taking place at the same time, prohibiting an extended experimental or computational investigation. In order to obtain more insight into this complex mechanism, models consisting of various elementary steps have been proposed [12,13]. Five classes of reversible reactions have been presented, including hydrogen abstraction, addition, substitution and cyclisation. The main concern is to establish the rate determining step for coke formation. In order to do so, *ab initio* modelling techniques are used to compute the energies of the involved molecules and the kinetics of the corresponding radical reactions. Additional information about the concentrations of the coke precursors

is obtained using single-event micro-kinetic simulation techniques, which in particular allows a comparison between the pseudo-first order (or modified) rate coefficients of the various elementary reactions [14]. Due to its reactive characteristics, the coke layer can be modelled as a network of aromatic rings. Initially, the coke growth process is modelled by investigating the route from benzene to naphthalene or naphthalene-like species [11]. In a next step, the influence of the polyaromatic environment is examined using an extended set of polycyclic aromatic hydrocarbons (PAHs) as model representatives [10]. Knowledge of the thermodynamic quantities of the latter species remains in general an ongoing challenge due to their high importance in a large variety of fields (such as incomplete combustion processes, astrophysics, full-erene chemistry...) [15,16]. In addition, PAHs also form the largest class of carcinogenic and mutagenic species [17,18].

Both thermodynamic and kinetic properties have been calculated using *ab initio* techniques. As the involved species are considerably large, density functional theory (DFT) is preferred over post-Hartree–Fock methods due to the interesting accuracy–computational cost ratio of the former theory. At present, the performance of the newest class of DFT functionals is known to be of sufficiently high quality leading to accurate and reliable data [19]. This aspect has in particular been shown for the essential

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class of hydrogen abstractions leading to aryl radicals by the presenting authors [10,20].

In the present contribution, various topics are being discussed. Firstly, DFT results corresponding to elementary reactions for a pathway starting from toluene and leading to 1-methylnaphthalene are given. So far, computational studies on methylated polyaromatics have been very limited [21]. The latter species are nevertheless known to play a key role within combustion methods. Abstraction of a methyl hydrogen leads to the class of benzylic radicals, in which the stability is enhanced due to a well-known resonance effect [22]. The influence of the polyaromatic environment is thus expected to be substantially larger compared to the series of aryl radicals. Secondly, the importance of benzylic radicals in growth processes of polyaromatics is highlighted further. An extended amount of bond dissociation enthalpies (BDEs) is computed in order to assess the reactivity of the species. Thirdly, radical hydrogen abstractions at a series of methylated polyaromatics using an assisting methyl radical are investigated.

## 2. Methodology

All calculations were performed using the Gaussian03 software package [23] (Wallingford, Connecticut 06492, US). The optimisation of geometries as well as frequency computations were performed using the B3LYP functional [24,25], in combination with a basis set of moderate size (6-31G + G(d,p) or 6-311G(d,p)). Based on other studies involving similar radical reactions, this method has proven to provide a reliable and quantitatively good description. For the computation of the thermodynamic and kinetic parameters of the studied hydrogen abstraction reactions, a factor of 0.9806, 0.9614 and 0.9989 was used to scale the zero-point vibrational energies (ZPVEs), the harmonic frequencies and the thermal correction to the enthalpy, respectively. The use of scale factors provides a mean for accounting for systematic deviations between measured and computed frequency-dependent properties, and is an important consideration for an accurate *ab initio* description [26]. Subsequent single-point energy computations were carried out applying the hybrid meta Boese–Martin for Kinetics (BMK) functional [27], in combination with the extended 6-311 + G(3df,2p) basis set. This particular functional was earlier shown to be accurate within approximately 10 kJ/mol for the calculation of reaction barriers [27]. Moreover, it provides precise thermodynamic and kinetic results for a variety of similar species and hydrogen abstraction reactions at a feasible computational cost [19,20,28]. We applied traditional transition state theory (TST) to obtain intrinsic rate coefficients for the hydrogen abstraction reactions at methylated PAHs. Special attention was given to the

accurate treatment of internal rotations of the methyl tops, present in both the reactants and the transition structures [29]. The accuracy of the results was increased by incorporating Eckart tunnelling corrections [30].

## 3. Results and discussion

### 3.1 Network starting from toluene

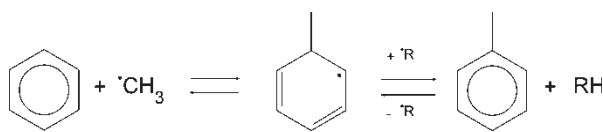
Compared to our earlier work on pure PAHs, we now start from a methylated species (toluene). The latter molecule is created following a substitution reaction at benzene (Scheme 1): an initial addition reaction of a methyl radical at benzene is followed by a hydrogen abstraction reaction. We earlier reported that this initial addition reaction of  $\text{CH}_3$  is highly competitive with the hydrogen abstraction reaction [11]. Toluene-like species can also be formed within the coke surface following other reactions, e.g. the addition of propene, and will inherently be present in the network.

A reaction route, starting from toluene (1-1) and leading to 1-methylnaphthalene (2-1) is presented in Scheme 2 (for the labelling of the species, see Figure 3). Analogous to earlier work, the pathway consists of consecutive elementary steps. The initial hydrogen abstraction reaction (R1) is followed by two additions of ethylene (R2 and R3). Two successive isomerisation reactions (R4 and R5) are observed, and are followed by a cyclisation reaction (R6). In order to restore the aromaticity of the newly formed ring, five dehydrogenation reactions are needed (R7–R11). Ultimately, 1-methylnaphthalene is formed.

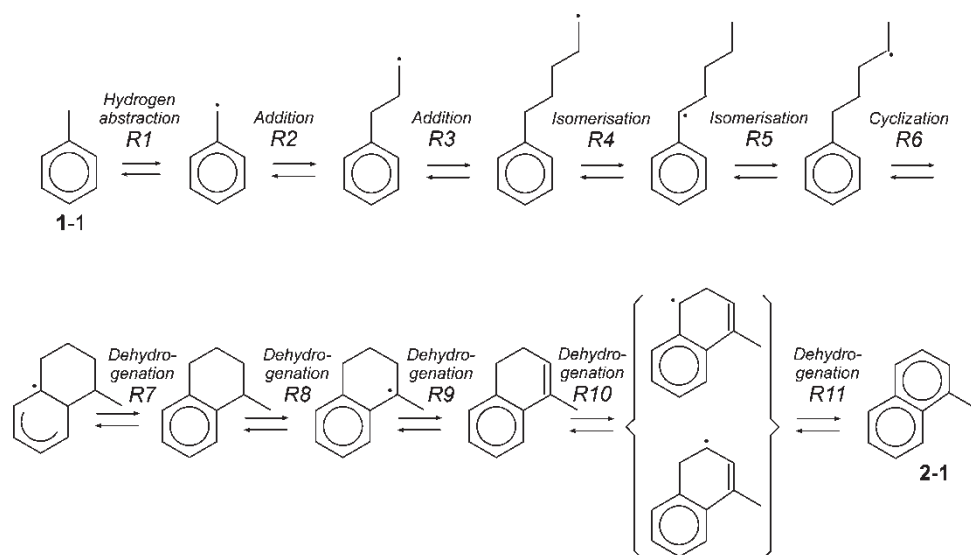
It is important to highlight that knowledge of the intrinsic rate coefficients  $k(T)$  is not sufficient in order to determine the rate limiting step of the entire pathway which contains both bimolecular and unimolecular reactions. The concentrations of the coke precursors have been taken into account. Pseudo first-order (or modified) rate coefficients  $k'(T)$  which are expressed in units 1/s are used:

$$r = k(T)c_Ac_B = k'(T)c_A.$$

The concentrations of the methyl radical, the ethylene molecule and the hydrogen radical were obtained from a two-dimensional reactor model developed at the Laboratory for Chemical Technology [14]. The radical coke precursors are present in much lower



Scheme 1. Substitution reaction starting from benzene and leading to toluene.



Scheme 2. Reaction pathway starting from toluene (**1-1**) and leading to 1-methylnaphthalene (**2-1**) (labelling see Figure 3).

concentrations compared to the olefinic species in the gas phase and this has a huge effect on the relative ratio of the modified rate coefficients. This is illustrated in Figure 1, where the intrinsic and modified rate coefficients are given for the initial addition reaction of a methyl radical at benzene (see Scheme 1). In the left parts of (a) and (b), the values obtained for the entire reactor tube are given. The  $x$ -axis corresponds to the relative radial coordinate with  $x = 0.0$  corresponding to the central axis and  $x = 1.0$  to the wall of the tube. The  $y$ -axis corresponds to the axial coordinate. In the right parts of (a) and (b), the part close to the wall of the tube where coke formation is initialised is given in more detail. The influence of the concentration of the methyl radical is enormous; the rate coefficient calculated at the tube wall for the reaction under consideration is reduced ( $k'(T)$  compared to  $k(T)$ ) with approximately a factor of 10,000.

We now return to the reaction network depicted in Scheme 2. An overview of the pseudo first-order rate coefficients  $k'(T)$  (in  $\text{s}^{-1}$ ) obtained near the wall of the reactor tube outlet is given in Figure 2. In analogy with earlier results (the reaction network from benzene to naphthalene [11]), the initial hydrogen abstraction at toluene (R1) and the final dehydrogenation reactions (going from a neutral to a radical species: R8 and R10) are found to be the crucial steps, as their respective  $k'(T)$  values are the smallest.

### 3.2 BDEs of benzylic radicals

In the previous paragraphs, we demonstrated the importance of toluene and the subsequently created benzyl radical as a starting point for the coke growth model. Moreover, we pointed out the importance of benzylic

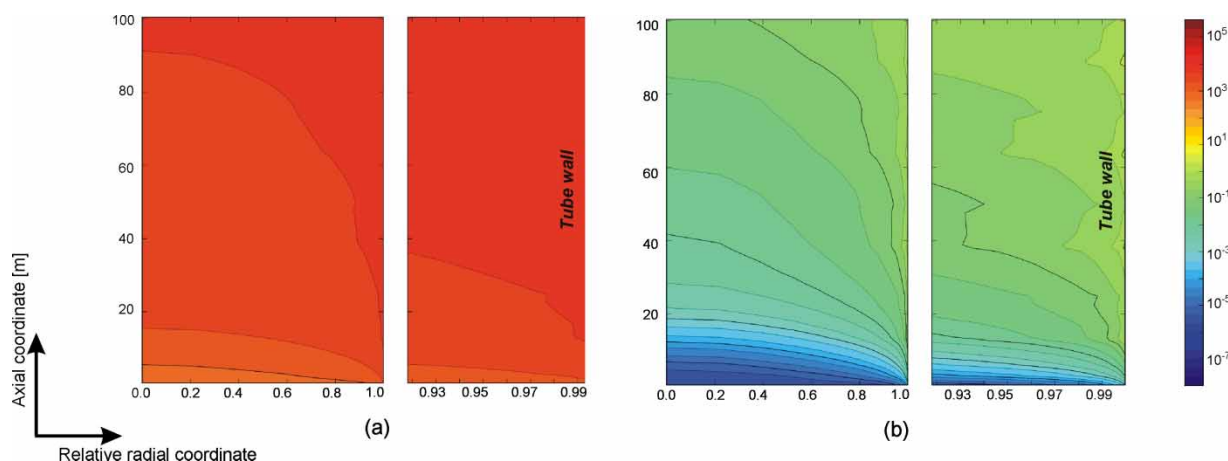


Figure 1. Intrinsic (a) and modified (b) rate coefficients for the addition reaction of a methyl radical to benzene (see Scheme 1).

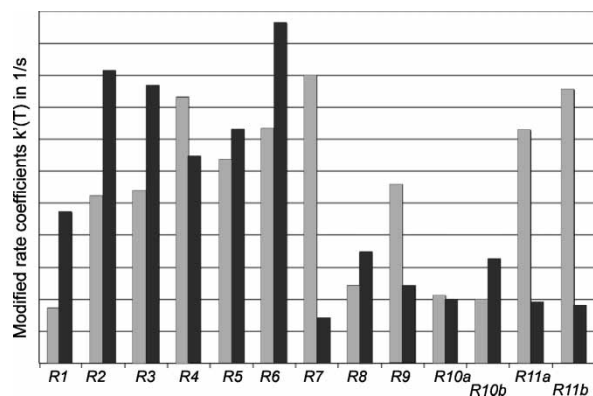


Figure 2. Modified rate coefficients  $k'(T)$  for the forward (in light grey) and reverse (in dark grey) reaction steps of the entire network (as defined in Scheme 2).

radicals for general growth processes of polyaromatics. In a next step, an extended set of singly methylated species, as depicted in Figure 3 is studied to account for the influence of the polyaromatic environment. The numbering of the various methylated sites conforms to the IUPAC convention.

More precisely, carbon–hydrogen BDEs, corresponding to the bond breaking of a methyl hydrogen, are calculated using the reliable and yet cost-effective

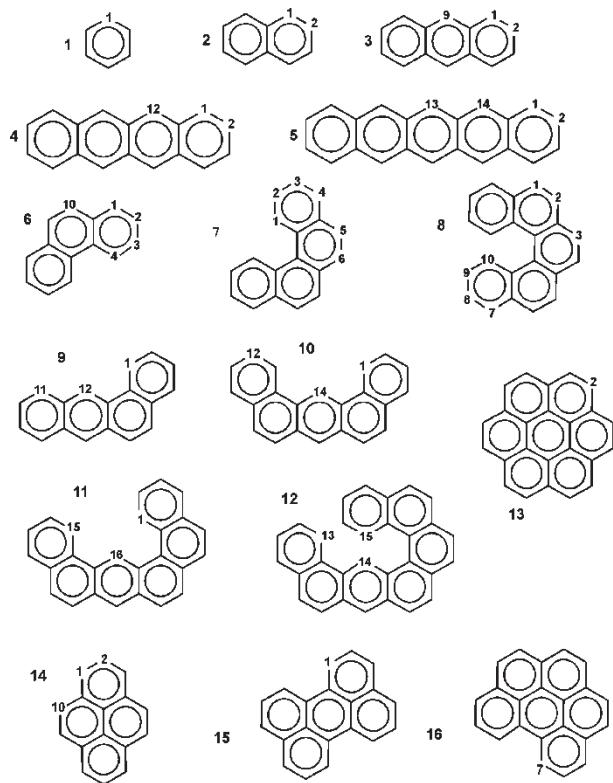


Figure 3. Singly methylated PAHs, methylation can occur at various positions as indicated (the methyl group is not displayed in the figure).

BMK/6-311 + G(3df,2p)//B3LYP/6-31 + G(d,p) level of theory. Izgorodina et al. reported on the potential problems using DFT functionals for reproducing reliable trends in BDE values as well as on the very reasonable performance of the BMK functional for the computation of a large set of absolute BDEs [31]. The results are given in Table 1. First of all we note that the BDE values of benzylic radicals are relatively small. This is in accordance with the enhanced stability of these species due to well-known resonance effects.

For the subcategory of non-linear species the effect of the polyaromatic environment is found to be smaller than expected. For example, the BDE value of the large methylated coronene molecule (structure **13-2**) deviates in a minor way (i.e. with 1.26 kJ/mol) from the reference BDE value of toluene (structure **1-1**). It is overall concluded that an enlarged polyaromatic environment (leading to an enhanced delocalisation possibility) does not lead to a substantially smaller BDE value. The behaviour of the series of linear methylated PAHs (structures **1–5**) is nevertheless very peculiar. Two distinct and equally important features can be distinguished. First of all, the BDE values do decrease with increasing molecular size for this important subcategory. The calculated BDE value of 301 kJ/mol for **5-13** deviates strongly from the BDE value at toluene (**1-1**, 373 kJ/mol), being the smallest representative of the linear series. Secondly, the BDE values strongly depend on the site of methylation, as is in particular seen for the structures **3**, **4** and **5**. This must be traced back to characteristic geometrical features – and in particular closely positioned hydrogen atoms introducing strong repulsive effects – in the optimised methylated reactant structures.

### 3.3 Hydrogen abstraction reactions at methylated PAHs

In the next step, abstraction reactions of methyl hydrogen at the proposed set of methylated PAHs (Figure 3) using an assisting methyl radical are considered. In this way, a benzylic radical and methane are formed. Thermodynamic and kinetic properties are computed, using BMK/6-311 + G(3df,2p) energies and B3LYP/6-31 + G(d,p) geometries. For the computation of intrinsic rate coefficients, Eckart tunnelling corrections are included in order to improve the accuracy of the standard TST results. It was however impossible to identify the majority of the internal rotations of the methyl tops (present in both the reactants and the transition states) as independent movements within the vibrational spectrum and therefore, all motions were treated using the harmonic oscillator model.

First of all it is noted that the geometrical features of the transition states indicate that the forming bond is always longer than the breaking bond, a behaviour



Table 1. BDE values of set of methylated PAHs.

PAH	BDE	PAH	BDE	PAH	BDE	PAH	BDE	PAH	BDE	PAH	BDE
<b>1-1</b>	372.83	<b>5-13</b>	301.29	<b>7-1</b>	365.12	<b>8-7</b>	369.94	<b>10-12</b>	373.67	<b>14-2</b>	374.24
<b>2-1</b>	368.53	<b>5-14</b>	315.11	<b>7-2</b>	369.00	<b>8-1</b>	367.36	<b>11-1</b>	366.00	<b>14-10</b>	368.70
<b>2-2</b>	370.00	<b>5-1</b>	355.63	<b>7-3</b>	370.77	<b>8-2</b>	370.71	<b>11-16</b>	341.51	<b>15-1</b>	351.50
<b>3-9</b>	345.83	<b>5-2</b>	351.56	<b>7-4</b>	370.33	<b>8-3</b>	369.77	<b>11-15</b>	374.92	<b>16-7</b>	358.59
<b>3-1</b>	363.44	<b>6-4</b>	368.75	<b>7-5</b>	369.05	<b>9-1</b>	369.94	<b>12-13</b>	371.92		
<b>3-2</b>	363.87	<b>6-3</b>	370.01	<b>7-6</b>	371.24	<b>9-12</b>	348.42	<b>12-14</b>	342.45		
<b>4-12</b>	328.08	<b>6-2</b>	371.61	<b>8-10</b>	367.68	<b>9-11</b>	364.43	<b>12-15</b>	367.21		
<b>4-1</b>	359.20	<b>6-1</b>	371.81	<b>8-9</b>	369.73	<b>10-14</b>	342.36	<b>13-2</b>	371.57		
<b>4-2</b>	357.39	<b>6-10</b>	369.98	<b>8-8</b>	370.16	<b>10-1</b>	369.08	<b>14-1</b>	361.64		

Note:  $\text{BDE}_{298}(R-H) = [\Delta_f H_{298}(R) + \Delta_f H_{298}(H)] - \Delta_f H_{298}(R-H)$  with  $\Delta_f H_{298}$  the enthalpies of formation of the radical species ( $R$ ) and molecule ( $R-H$ ), respectively. The following nomenclature (referring to Figure 3) is used: The bold number corresponds to the compound, whereas the second number refers to the isomeric position. Values calculated using B3LYP/6-31 + G(d,p) geometries and BMK/6-311 + G(3df,2p) scf energies.

which is in correspondence with the enhanced stability of the product radicals: the transition states can be described as ‘early’ in the Hammond sense, as they geometrically resemble the reactants. Opposite to the series of pure (or non-methylated) PAHs, the influence of the planarity of the aromatic substrate is less substantial. Moreover, in case of the transition structures the methylation usually results in a non-planar structure in which the attached methyl group is twisted away from the aromatic substrate. Consequently, the attacking gas phase radical (in this particular case a methyl radical) encounters relatively little or no steric hindrance when approaching the site of abstraction.

Secondly, the thermodynamic properties are calculated using scaled ZPVEs and thermal corrections to the enthalpy. The reaction barriers at 0 K ( $\Delta E_0$ ) and the reaction enthalpies at 298 K ( $\Delta H_{298}$ ) reveal interesting characteristics. The latter values are of course closely related to the BDE results which were discussed in section 3.2. For the present contribution, it is important to address the reactivity–enthalpy relationship. We earlier found that for the series of abstraction at non-methylated PAHs (leading to the formation of aryl radicals), a normal reactivity–enthalpy relationship could only be found for hydrogen abstractions at the subcategory of linear acenes as in these cases, steric hindrance effects between the aromatic substrate on one hand and the incoming methyl radical on the other, are non-existent [10]. For the present species we mentioned earlier that steric effects are not important for the majority of the transition structures. And indeed, a normal reactivity–enthalpy was obtained overall: A large stabilisation effect in the product radical (large absolute  $\Delta H_{298}$  value) corresponds to an enhanced reactivity (small  $\Delta E_0$  value). This behaviour is depicted in Figure 4. Several exceptional points can, however, be discerned within the figure. They correspond to reactions with barriers around (or larger than) 50 kJ/mol and they mostly correspond to strongly congested sites (such as **11-16** and **12-14**).

Thirdly, the rate coefficients  $k(T)$  and derived kinetic properties (activation energy  $E_a$  and pre-exponential factor  $A$ ) are computed for the temperature interval between 700 and 1100 K, which is relevant for the coke growth process. It is noted that, compared to abstractions at pure PAHs, the majority of the abstractions at methylated PAHs occurs faster. This is illustrated in Figure 5, for a random selection of five reactions. Besides the abstraction at the reference molecule (benzene and toluene, respectively), 2 abstractions at linear species (in particular **3-9** and **5-2**) and 2 abstractions at non-linear species (**10-14** and **11-1**) are selected. It is indeed seen that for the majority of the cases, the abstraction at methylated species are faster in comparison with abstractions at non-methylated PAHs. Although it is hard to draw conclusions for the entire test set, it is in general found that abstraction is easier at larger systems. These results underline the importance of methylated species and the corresponding hydrogen abstraction reactions to initiate radical sites within a polyaromatic surface.

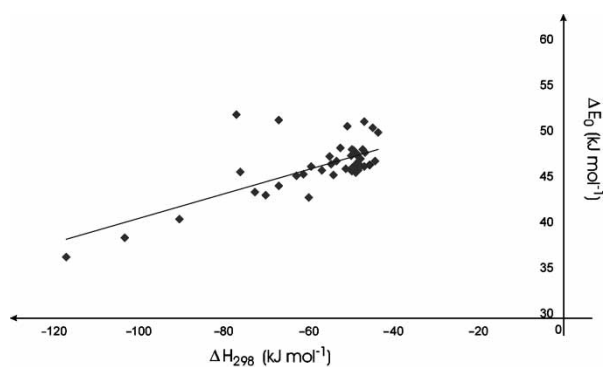


Figure 4. Reactivity–enthalpy relationship for the hydrogen abstraction reactions at methylated PAHs by an attacking methyl radical.

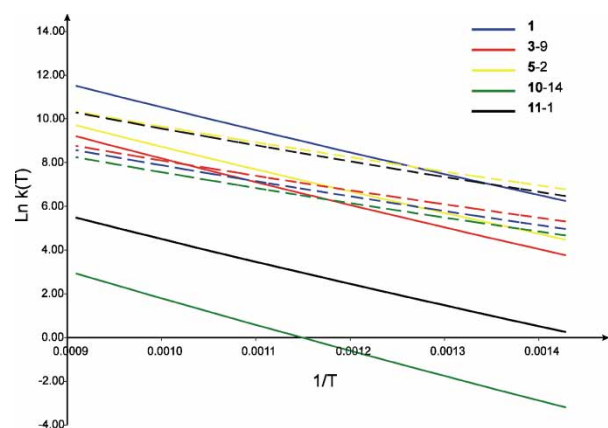


Figure 5. Rate constants  $\ln k(T)$  ( $k(T)$  in  $\text{m}^3/\text{mol/s}$ ) as a function of  $1/T$  (in  $1/K$ ) for a random selection of five hydrogen abstraction reactions (numbering conforms with Figure 3). The full lines correspond to abstractions at pure PAHs (leading to aryl radicals) whereas the dotted lines correspond to abstractions at methylated PAHs (leading to benzylic radicals).

#### 4. Conclusions

In this contribution we provided an overview of recent results obtained within the *ab initio* study of radical reactions leading to the growth of polyaromatic species. The examined growth process can be situated within the industrially important application of coke formation at the inner wall of a reactor tube used in a steam cracking procedure. The latter remains at present a central production methodology of light olefins such as ethylene and propylene.

First of all, a new reaction route starting from toluene and leading to 1-methylnaphthalene has been proposed. A comparison between the pseudo first-order rate coefficients of all elementary reactions indicates that the initial hydrogen abstraction reaction at toluene (leading to the formation of the benzyl radical) and the final dehydrogenation reactions are the rate limiting steps. The importance of benzylic radicals was also highlighted. Consequently, carbon–hydrogen bond dissociation enthalpies (BDEs) were calculated for an extended set of methylated PAHs, being model representatives for the coke surface. The effect of the polyaromatic environment on BDEs is found to be smaller than expected for the subcategory of non-linear species. The subcategory of linear methylated PAHs on the other hand reveals peculiar behaviour. First of all, the BDE values decrease with increasing molecular size. And secondly, the BDE values depend in a very substantial way on the site of methylation. Finally, hydrogen abstraction reactions at the set of methylated PAHs using an assisting methyl radical were closely investigated. The abstraction is found to preferentially occur at the larger systems and is in general faster compared to abstractions at the analogous non-methylated species. In this contribution, the presented

overview of reliable DFT-results on methylated species and derived benzylic radicals underlines their importance within general growth schemes of polyaromatic structures. The corresponding hydrogen abstraction reactions are found to be key reactions to initiate radical sites within the polyaromatic surface.

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